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Note to students: Whenever balancing an equation, it is acceptable to leave spaces blank instead of writing “1’ – in chemistry, they mean the same thing. 1) 1 AlBr3 + 3 K 3 KBr + 1 Al 2) 1 FeO + 1 PdF2 1 FeF2 + 1 PdO 3) 1 P4 + 6 Br2 4 PBr3 4) 2 LiCl + 1 Br2 2 LiBr + 1 Cl2 5) 1 PbBr2 + 2 HCl 2 HBr + 1 PbCl2 6) 2 CoBr3 + 3 CaSO4 3 CaBr2 + 1 Co2(SO4)3 7) 2 Na3P + 3 CaF2 6 NaF + 1 Ca3P2 8) 2 Mn + 6 HI 3 H2 + 2 MnI3 9) 1 Li3PO4 + 3 NaBr 1 Na3PO4 + 3 LiBr 10) 1 CaF2 + 1 Li2SO4 1 CaSO4 + 2 LiF 11) 2 HBr + 1 Mg(OH)2 1 MgBr2 + 2 H2O 12) 2 LiNO3 + 1 CaBr2 1 Ca(NO3)2 + 2 LiBr 13) 1 AgNO3 + 1 Li 1 LiNO3 + 1 Ag 14) 1 Si(OH)4 + 4 NaBr 1 SiBr4 + 4 NaOH 15) 2 NaCN + 1 CuCO3 1 Na2CO3 + 1 Cu(CN)2 © 2004 Cavalcade Publishing, All Rights Reserved For chemistry help, visit www.chemfiesta.com Activate your 30 day free trial to unlock unlimited reading. 1. Another Balancing Equations Sheet! 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Then go back and balance the following equations: To practice balancing, you may use the Phet Lab online. When finished, check your answers. 1) ___ NaBr + ___ H3PO4 ___ Na3PO4 + ___ HBr 2) ___ Ca(OH)2 + ___ Al2(SO4)3 ___ CaSO4 + ___ Al(OH)3 Type of reaction: _____ 3) ___ Mg + ___ Fe2O3 ___ Fe + ___ MgO Type of reaction: _____ 4) ___ C2H4 + ___ O2 ___ CO2 + ___ H2O Type of reaction: _____ 5) ___ PbSO4 ___ PbSO3 + ___ O2 Type of reaction: _____ 6) ___ NH3 + ___ I2 ___ N2I6 + ___ H2 Type of reaction: _____ 7) ___ H2O + ___ SO3 ___ H2SO4 Type of reaction: _____ 8) ___ H2SO4 + ___ NH4OH ___ H2O + ___ (NH4)2SO4 Type of reaction: _____ Type of reaction: _____ Balancing Equations Practice Worksheet Balance the following equations: 1) ___ NaNO3 + ___ PbO ___ Pb(NO3)2 + ___ Na2O 2) ___ AgI + ___ Fe2(CO3)3 ___ FeI3 + ___ Ag2CO3 3) ___ C2H4O2 + ___ O2 ___ CO2 + ___ H2O 4) ___ ZnSO4 + ___ Li2CO3 ___ ZnCO3 + ___ Li2SO4 5) ___ V2O5 + ___ CaS ___ CaO + ___ V2S5 6) ___ Mn(NO2)2 + ___ BeCl2 ___ Be(NO2)2 + ___ MnCl2 7) ___ AgBr + ___ GaPO4 ___ Ag3PO4 + ___ GaBr3 8) ___ H2SO4 + ___ B(OH)3 ___ B2(SO4)3 + ___ H2O 9) ___ S8 + ___ O2 ___ SO2 10) ___ Fe + ___ AgNO3 ___ Fe(NO3)2 + ___ Ag Another Balancing Equations Sheet! Balance these equations! 1) ___ AlBr3 + ___ K ___ KBr + ___ Al 2) ___ FeO + ___ PdF2 ___ FeF2 + ___ PdO 3) ___ P4 + ___ Br2 ___ PBr3 4) ___ LiCl + ___ Br2 ___ LiBr + ___ Cl2 5) ___ PbBr2 + ___ HCl ___ HBr + ___ PbCl2 6) ___ CoBr3 + ___ CaSO4 ___ CaBr2 + ___ Co2(SO4)3 7) ___ Na3P + ___ CaF2 ___ NaF + ___ Ca3P2 8) ___ Mn + ___ HI ___ H2 + ___ MnI3 9) ___ Li3PO4 + ___ NaBr ___ Na3PO4 + ___ LiBr 10) ___ CaF2 + ___ Li2SO4 ___ CaSO4 + ___ LiF 11) ___ HBr + ___ Mg(OH)2 ___ MgBr2 + ___ H2O 12) ___ LiNO3 + ___ CaBr2 ___ Ca(NO3)2 + ___ LiBr 13) ___ AgNO3 + ___ Li ___ LiNO3 + ___ Ag 14) ___ Si(OH)4 + ___ NaBr ___ SiBr4 + ___ NaOH 15) ___ NaCN + ___ CuCO3 ___ Na2CO3 + ___ Cu(CN)2 Word Equations Worksheet Write and balance the following chemical equations. 1) When dissolved beryllium chloride reacts with dissolved silver nitrate in water, aqueous beryllium nitrate and silver chloride powder are made. 2) When isopropanol (C3H8O) burns in oxygen, carbon dioxide, water, and heat are produced. 3) When dissolved sodium hydroxide reacts with sulfuric acid (H 2SO4), aqueous sodium sulfate, water, and heat are formed. 4) When fluorine gas is put into contact with calcium metal at high temperatures, calcium fluoride powder is created in an exothermic reaction. 5) When sodium metal reacts with iron (II) chloride, iron metal and sodium chloride are formed. Types of Reactions Worksheet – Solutions Balance the following equations and indicate the type of reaction taking place: 1) 3 NaBr + 1 H3PO4 1 Na3PO4 + 3 HBr Type of reaction: double displacement 2) 3 Ca(OH)2 + 1 Al2(SO4)3 3 CaSO4 + 2 Al(OH)3 Type of reaction: double displacement 3) 3 Mg + 1 Fe2O3 2 Fe + 3 MgO Type of reaction: single displacement 4) 1 C2H4 + 3 O2 2 CO2 + 2 H2O Type of reaction: combustion 5) 2 PbSO4 2 PbSO3 + 1 O2 Type of reaction: decomposition 6) 2 NH3 + 3 I2 1 N2I6 + 3 H2 Type of reaction: double displacement 7) 1 H2O + 1 SO3 1 H2SO4 Type of reaction: decomposition 8) 1 H2SO4 + 2 NH4OH 2 H2O + 1 (NH4)2SO4 Type of reaction: acid-base Solutions for the Balancing Equations Practice Worksheet 1) 2 NaNO3 + PbO Pb(NO3)2 + Na2O 2) 6 AgI + Fe2(CO3)3 2 FeI3 + 3 Ag2CO3 3) C2H4O2 + 2 O2 2 CO2 + 2 H2O 4) ZnSO4 + Li2CO3 ZnCO3 + Li2SO4 5) V2O5 + 5 CaS 5 CaO + V2S5 6) Mn(NO2)2 + BeCl2 Be(NO2)2 + MnCl2 7) 3 AgBr + GaPO4 Ag3PO4 + GaBr3 8) 3 H2SO4 + 2 B(OH)3 B2(SO4)3 + 6 H2O 9) S8 + 8 O2 8 SO2 10) Fe + 2 AgNO3 Fe(NO3)2 + 2 Ag Another Balancing Equations Sheet! – Answers Balance these equations! Note to students: Whenever balancing an equation, it is acceptable to leave spaces blank instead of writing “1’ – in chemistry, they mean the same thing. 1) 1 AlBr3 + 3 K 3 KBr + 1 Al 2) 1 FeO + 1 PdF2 1 FeF2 + 1 PdO 3) 1 P4 + 6 Br2 4 PBr3 4) 2 LiCl + 1 Br2 2 LiBr + 1 Cl2 5) 1 PbBr2 + 2 HCl 2 HBr + 1 PbCl2 6) 2 CoBr3 + 3 CaSO4 3 CaBr2 + 1 Co2(SO4)3 7) 2 Na3P + 3 CaF2 6 NaF + 1 Ca3P2 8) 2 Mn + 6 HI 3 H2 + 2 MnI3 9) 1 Li3PO4 + 3 NaBr 1 Na3PO4 + 3 LiBr 10) 1 CaF2 + 1 Li2SO4 1 CaSO4 + 2 LiF 11) 2 HBr + 1 Mg(OH)2 1 MgBr2 + 2 H2O 12) 2 LiNO3 + 1 CaBr2 1 Ca(NO3)2 + 2 LiBr 13) 1 AgNO3 + 1 Li 1 LiNO3 + 1 Ag 14) 1 Si(OH)4 + 4 NaBr 1 SiBr4 + 4 NaOH 15) 2 NaCN + 1 CuCO3 1 Na2CO3 + 1 Cu(CN)2 Word Equations Worksheet - Solutions Write the word equations for each of the following chemical reactions: 1) 2) 3) 4) 5) BeCl2(aq) + 2 AgNO3(aq) Be(NO3)2(aq) + 2 AgCl(s) 2 C3H8O(l) + 9 O2(g) 6 CO2(g) + 8 H2O(g) 2 NaOH(aq) + H2SO4(l) Na2SO4 + 2 H2O(l) F2(g) + Ca(s) CaF2(s) 2 Na(s) + FeCl2(s) 2 NaCl(s) + Fe(s)

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